

- Indicate the number of each in the following:
 - Orbitals in the 3d sublevel **5 orbitals**
 - Sublevels in the $n = 1$ principal level **1 sublevel (s)**
 - Orbitals in the 6s sublevel **1 orbital**
 - Sublevels in the $n = 3$ principal level **3 sublevels (s, p, and d)**
- Indicate the type and number of **orbitals** in each of the following principle energy levels or sublevels:
 - 3p sublevel **Three p orbitals**
 - $n = 2$ **One s and Three p orbitals**
 - 4d sublevel **Five d Orbitals**
 - 1s sublevel **One s Orbital**
 - $n = 3$ **1 (s) / 3 (p) / 5 (d) Orbitals**
 - $n = 4$ **1 (s) / 3 (p) / 5 (d) / 7 (f) Orbitals**
 - 5f sublevel **Seven f Orbitals**
- Indicate the maximum number of **electrons** in each of the following:
 - 2p orbital **2 electrons**
 - 3p sublevel **6 electrons**
 - $n = 4$ principal level **32 electrons**
 - 5d sublevel **10 electrons**
 - 3s sublevel **2 electrons**
 - $n = 3$ principal level **18 electrons**
 - 4p orbital **2 electrons**
 - 5f sublevel **14 electrons**