

1. Indicate the number of each in the following:

- A) Orbitals in the 3d sublevel \_\_\_\_\_
- B) Sublevels in the  $n = 1$  principal level \_\_\_\_\_
- C) Orbitals in the 6s sublevel \_\_\_\_\_
- D) Sublevels in the  $n = 3$  principal level \_\_\_\_\_

2. Indicate the type and number of **orbitals** in each of the following principle energy levels or sublevels:

- A) 3p sublevel \_\_\_\_\_
- B)  $n = 2$  \_\_\_\_\_
- C) 4d sublevel \_\_\_\_\_
- D) 1s sublevel \_\_\_\_\_
- E)  $n = 3$  \_\_\_\_\_
- F)  $n = 4$  \_\_\_\_\_
- G) 5f sublevel \_\_\_\_\_

3. Indicate the maximum number of **electrons** in each of the following:

- A) 2p orbital \_\_\_\_\_
- B) 3p sublevel \_\_\_\_\_
- C)  $n = 4$  principal level \_\_\_\_\_
- D) 5d sublevel \_\_\_\_\_
- E) 3s sublevel \_\_\_\_\_
- F)  $n = 3$  principal level \_\_\_\_\_
- G) 4p orbital \_\_\_\_\_
- H) 5f sublevel \_\_\_\_\_