

You walk into the Gen Chem lab to find a beaker filled with 350.0 mL of 0.325 M dimethylamine ((CH<sub>3</sub>)<sub>2</sub>NH; K<sub>b</sub> = 5.4 × 10<sup>-4</sup>). Above the beaker is a buret filled with 0.750 M HNO<sub>3</sub>. Your job is to titrate the (CH<sub>3</sub>)<sub>2</sub>NH with four additions of nitric acid and determine the pH. At least one of the additions must be the equivalence point.

In class today, we looked at 0 mL added, 152 mL added, 42 mL added and 200 mL added